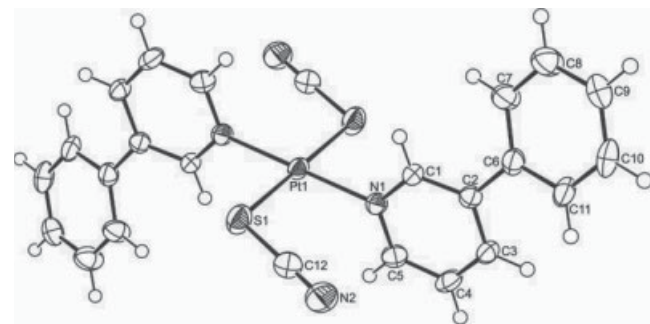


Crystal structure of bis(3-phenylpyridine- κN)bis(thiocyanato- κS)-platinum(II), $C_{24}H_{18}N_4PtS_2$

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Abstract

$C_{24}H_{18}N_4PtS_2$, monoclinic, $P2_1/n$ (no. 14), $a = 9.820(1)$ Å, $b = 5.3464(8)$ Å, $c = 21.463(3)$ Å, $\beta = 102.471(3)^\circ$, $V = 1100.3$ Å³, $Z = 2$, $R_{gt}(F) = 0.0254$, $wR_{ref}(F^2) = 0.0578$, $T = 200$ K.

Table 1. Data collection and handling.

Crystal:	yellow sticks, size 0.06×0.09×0.34 mm
Wavelength:	Mo K_α radiation (0.71073 Å)
μ :	65.86 cm ⁻¹
Diffractometer, scan mode:	Bruker SMART 1000 CCD, φ and ω
$2\theta_{max}$:	52.08°
$N(hkl)_{measured}$, $N(hkl)_{unique}$:	6373, 2138
Criterion for I_{obs} , $N(hkl)_{gt}$:	$I_{obs} > 2 \sigma(I_{obs})$, 1534
$N(param)_{refined}$:	142
Programs:	SHELX [5], ORTEP-3[6], PLATON [7]

Source of material

To a solution of K_2PtCl_4 (0.205 g, 0.493 mmol) in H_2O (40 ml) were added KSCN (0.505 g, 5.194 mmol) and 3-phenylpyridine (*ppy*; 0.174 g, 1.121 mmol) in EtOH (10 ml) and stirred for 24 h at room temperature. The formed precipitate was separated by filtration, washed with H_2O and EtOH, and dried at 50 °C, to give a pale yellow powder (0.244 g). Crystals suitable for X-ray diffraction analysis were obtained by slow evaporation from a CH_3CN solution.

Experimental details

Hydrogen atoms were positioned geometrically and allowed to ride on their parent atoms with $d(C-H) = 0.95$ Å and $U_{iso}(H) = 1.2U_{eq}(C)$. The highest peak (1.16 e⁻Å⁻³) and the deepest hole (−0.90 e⁻Å⁻³) in the difference Fourier map are located 0.89 Å and 0.89 Å from the atom Pt1, respectively.

Discussion

Crystal structures of the related Pd(II) complex $[PdCl_2(ppy)_2]$ and Pt(IV) anionic complex $(H-ppy)[PtCl_5(ppy)] \cdot H_2O$ have been reported previously [1, 2]. The central Pt(II) ion in the title complex $[Pt(SCN)_2(ppy)_2]$ has a *trans*-S₂N₂ square-planar coordination

defined by two N atoms from two 3-phenylpyridine ligands and two S atoms of two SCN[−] anions. The Pt atom is located on an inversion center, and therefore the asymmetric unit contains one half of the complex and the PtS₂N₂ moiety is exactly planar. The pyridine ring of the ligand is almost perpendicular to the PtS₂N₂ plane (86.4(1)°). In the crystal structure, the *ppy* ligand is not planar. The pyridine and phenyl rings are twisted by 20.7(2)°. The thiocyanato ligand is almost linear with a S–C–N bond angle of 179.0(5)°, and the S atoms are coordinated to the Pt atom with the nearly tetrahedral Pt–S–C bond angle of 105.4(2)°, characteristic of an S-bonded conformation [3, 4]. The complex molecules are stacked in columns along [010] with $d(Pt \cdots Pt) = 5.3464(8)$ Å (= length of *b* axis) and successive complexes stack in opposite directions along [001]. In the columns, several intermolecular π – π interactions between the aromatic rings are present. The shortest distance between Cg1 (the centroid of ring N1–C5) and Cg2ⁱ (the centroid of ring C6–C11, symmetry code *i*: *x*, *y*–1, *z*) is 3.777(3) Å, and the dihedral angle between the ring planes is 20.7(2)°. Moreover, the complex displays intermolecular C–H $\cdots$ N hydrogen bonds with $d(C \cdots N) = 3.395(7)$ Å and 3.402(7) Å.

Table 2. Atomic coordinates and displacement parameters (in Å²).

Atom	Site	<i>x</i>	<i>y</i>	<i>z</i>	<i>U</i> _{iso}
H(1)	4e	0.2448	0.3261	0.0503	0.031
H(3)	4e	0.4380	0.0680	0.2228	0.035
H(4)	4e	0.2665	−0.2358	0.2153	0.039
H(5)	4e	0.0895	−0.2586	0.1249	0.038
H(7)	4e	0.4377	0.4943	0.0448	0.048
H(8)	4e	0.6018	0.8124	0.0510	0.054
H(9)	4e	0.7373	0.9323	0.1489	0.047
H(10)	4e	0.7024	0.7405	0.2421	0.051
H(11)	4e	0.5362	0.4295	0.2364	0.041

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Table 3. Atomic coordinates and displacement parameters (in Å²).

Atom	Site	<i>x</i>	<i>y</i>	<i>z</i>	<i>U</i> ₁₁	<i>U</i> ₂₂	<i>U</i> ₃₃	<i>U</i> ₁₂	<i>U</i> ₁₃	<i>U</i> ₂₃
Pt(1)	2 <i>a</i>	0	0	0	0.0265(2)	0.0286(1)	0.0209(1)	−0.0043(1)	0.00010(9)	0.0005(1)
S(1)	4 <i>e</i>	−0.1587(1)	0.2408(3)	0.03968(6)	0.0373(8)	0.0517(8)	0.0301(7)	0.0068(7)	0.0001(5)	−0.0054(6)
N(1)	4 <i>e</i>	0.1528(4)	0.0307(7)	0.0791(2)	0.025(2)	0.032(2)	0.021(2)	−0.005(2)	0.002(2)	−0.002(2)
N(2)	4 <i>e</i>	−0.0238(5)	0.3692(9)	0.1653(2)	0.050(3)	0.060(3)	0.035(3)	−0.005(3)	0.008(2)	−0.009(2)
C(1)	4 <i>e</i>	0.2509(5)	0.2097(9)	0.0843(2)	0.024(3)	0.031(3)	0.021(2)	0.001(2)	0.003(2)	0.003(2)
C(2)	4 <i>e</i>	0.3598(5)	0.2318(8)	0.1365(2)	0.025(3)	0.027(2)	0.024(2)	0.003(2)	0.006(2)	−0.002(2)
C(3)	4 <i>e</i>	0.3645(5)	0.0600(8)	0.1859(2)	0.029(3)	0.037(3)	0.019(2)	0.003(2)	0.001(2)	0.000(2)
C(4)	4 <i>e</i>	0.2637(5)	−0.1202(9)	0.1814(2)	0.039(3)	0.034(3)	0.024(3)	0.006(3)	0.004(2)	0.009(2)
C(5)	4 <i>e</i>	0.1590(5)	−0.1330(9)	0.1279(2)	0.034(3)	0.030(3)	0.034(3)	0.000(2)	0.009(2)	0.007(2)
C(6)	4 <i>e</i>	0.4673(5)	0.4302(9)	0.1402(2)	0.025(3)	0.038(3)	0.026(3)	0.003(2)	0.002(2)	−0.005(2)
C(7)	4 <i>e</i>	0.4903(6)	0.5451(9)	0.0853(2)	0.038(3)	0.056(4)	0.026(3)	−0.016(3)	0.009(2)	−0.010(2)
C(8)	4 <i>e</i>	0.5891(6)	0.733(1)	0.0889(2)	0.047(4)	0.054(3)	0.037(3)	−0.016(3)	0.014(3)	−0.009(3)
C(9)	4 <i>e</i>	0.6688(5)	0.805(1)	0.1467(2)	0.027(3)	0.039(3)	0.052(3)	−0.006(2)	0.006(2)	−0.005(3)
C(10)	4 <i>e</i>	0.6481(5)	0.691(1)	0.2017(2)	0.030(3)	0.045(3)	0.043(3)	0.000(3)	−0.013(2)	−0.009(3)
C(11)	4 <i>e</i>	0.5490(5)	0.5066(9)	0.1982(2)	0.030(3)	0.037(3)	0.030(3)	0.001(3)	−0.007(2)	0.004(3)
C(12)	4 <i>e</i>	−0.0774(5)	0.3163(9)	0.1143(2)	0.039(3)	0.034(3)	0.035(3)	0.000(2)	0.010(2)	0.001(2)

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