

Craig D. Grimmer\*

# Crystal structure of phenyl(2-phenyl-2,3-dihydro-1*H*-perimidin-2-yl)methanone, C<sub>24</sub>H<sub>18</sub>N<sub>2</sub>O

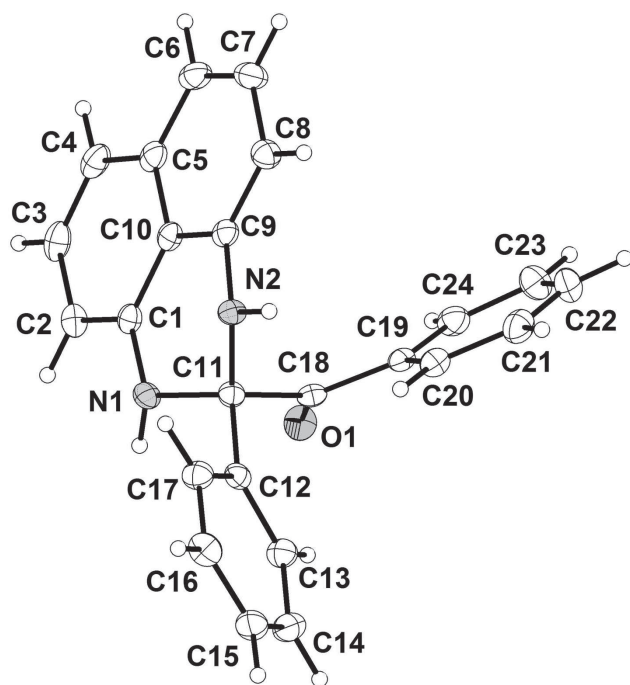


Table 1: Data collection and handling.

|                                                                         |                                                           |
|-------------------------------------------------------------------------|-----------------------------------------------------------|
| Crystal:                                                                | Yellow needle                                             |
| Size:                                                                   | 0.20 × 0.10 × 0.05 mm                                     |
| Wavelength:                                                             | Mo K $\alpha$ radiation (0.71073 Å)                       |
| $\mu$ :                                                                 | 0.8 cm <sup>-1</sup>                                      |
| Diffractometer, scan mode:                                              | Bruker APEX-II, $\varphi$ and $\omega$                    |
| 2 $\theta$ <sub>max</sub> , completeness:                               | 53°, >99%                                                 |
| $N(hkl)$ <sub>measured</sub> , $N(hkl)$ <sub>unique</sub> , $R_{int}$ : | 63640, 3704, 0.034                                        |
| Criterion for $I_{obs}$ , $N(hkl)$ <sub>gt</sub> :                      | $I_{obs} > 2 \sigma(I_{obs})$ , 3277                      |
| $N(param)$ <sub>refined</sub> :                                         | 247                                                       |
| Programs:                                                               | Bruker programs [1], SHELX [2], OLEX2 [3, 4], publCIF [5] |

## Source of material

Benzil (2.10 g; 10.0 mmol) and 1,8-diaminonaphthalene (1.58 g; 10.0 mmol) were dissolved in ethanol (75 mL). Concentrated HCl (aq) (4 drops) was added. The clear solution was heated to reflux for ~40 h during which a yellow precipitate formed. After cooling to room temperature, the yellow precipitate was removed by suction filtration, washed with ethanol (50 mL) and air-dried for ~20 h (1.94 g; 55%). The product was recrystallized from ethanol.

## Experimental details

All hydrogen atoms were placed in calculated positions and refined using a riding model. Carbon-bound H atoms were fixed at 0.93 Å, with displacement parameters 1.2 times  $U_{eq}(C)$ . Nitrogen-bound H atoms were fixed at 0.86 Å, with  $U_{iso}(H)$  set to 1.2 $U_{eq}(N)$ .

## Comment

Perimidines are an important class of compounds with attractive biological and physical properties [6–12]. The title compound [13–16] was prepared by condensation of 1,8-diaminonaphthalene and benzil with acid catalysis. One of the nitrogen atoms (N1) is coplanar with the ten carbon atoms of the naphthylene ring (C1–C10). The second nitrogen atom (N2) is slightly displaced below this plane by 0.14 Å and the bridging carbon atom (C11) is displaced above this plane by 0.56 Å. The plane of the phenyl ring (C12–C17) intersects with the naphthyl plane (C1–C10) at an angle of 67.9° and the

DOI 10.1515/ncrs-2016-0242

Received August 8, 2016; accepted January 18, 2017; available online February 25, 2017

## Abstract

C<sub>24</sub>H<sub>18</sub>N<sub>2</sub>O, monoclinic,  $P2_1/n$  (no. 14),  $a = 5.9227(7)$  Å,  $b = 16.462(2)$  Å,  $c = 18.150(2)$  Å,  $\beta = 98.372(2)^\circ$ ,  $V = 1750.8(4)$  Å<sup>3</sup>,  $Z = 4$ ,  $R_{gt}(F) = 0.0353$ ,  $wR_{ref}(F^2) = 0.0942$ ,  $T = 100(2)$  K.

CCDC no.: 1528291

The asymmetric unit of the title crystal structure is shown in the figure. Tables 1 and 2 contain details of the measurement method and a list of the atoms including atomic coordinates and displacement parameters.

\*Corresponding author: Craig D. Grimmer, School of Chemistry and Physics, University of Kwazulu-Natal, Pietermaritzburg 3201, South Africa, e-mail: craig.grimmer@gmail.com

**Table 2:** Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters (Å<sup>2</sup>).

| Atom | x           | y          | z          | <i>U</i> <sub>iso</sub> <sup>*</sup> / <i>U</i> <sub>eq</sub> |
|------|-------------|------------|------------|---------------------------------------------------------------|
| O1   | 0.03471(14) | 0.48645(5) | 0.21651(5) | 0.0207(2)                                                     |
| N2   | 0.55029(16) | 0.44201(6) | 0.15019(6) | 0.0143(2)                                                     |
| C19  | 0.38909(19) | 0.47854(7) | 0.29437(6) | 0.0151(2)                                                     |
| C5   | 0.2403(2)   | 0.24648(7) | 0.12042(6) | 0.0181(2)                                                     |
| N1   | 0.18919(16) | 0.46976(6) | 0.08326(5) | 0.0155(2)                                                     |
| H1   | 0.06885(16) | 0.50168(6) | 0.07585(5) | 0.0185(2)*                                                    |
| C8   | 0.6578(2)   | 0.30623(7) | 0.19638(7) | 0.0188(2)                                                     |
| H8   | 0.7984(2)   | 0.32563(7) | 0.22228(7) | 0.0226(3)*                                                    |
| C9   | 0.50445(19) | 0.35953(7) | 0.15790(6) | 0.0152(2)                                                     |
| C1   | 0.12935(19) | 0.38666(7) | 0.08258(6) | 0.0154(2)                                                     |
| C3   | −0.1226(2)  | 0.27446(8) | 0.04448(7) | 0.0221(3)                                                     |
| H3   | −0.2632(2)  | 0.25544(8) | 0.01822(7) | 0.0265(3)*                                                    |
| C2   | −0.0746(2)  | 0.35823(8) | 0.04514(6) | 0.0187(2)                                                     |
| H2a  | −0.1825(2)  | 0.39510(8) | 0.01984(6) | 0.0224(3)*                                                    |
| C4   | 0.0290(2)   | 0.21959(8) | 0.08091(7) | 0.0212(3)                                                     |
| H4   | −0.0077(2)  | 0.16339(8) | 0.07965(7) | 0.0254(3)*                                                    |
| C12  | 0.42314(19) | 0.58393(7) | 0.13764(6) | 0.0148(2)                                                     |
| C10  | 0.29050(19) | 0.33105(7) | 0.12067(6) | 0.0155(2)                                                     |
| C18  | 0.24208(19) | 0.48786(7) | 0.22045(6) | 0.0145(2)                                                     |
| C13  | 0.3118(2)   | 0.64663(7) | 0.16946(7) | 0.0185(2)                                                     |
| H13  | 0.1904(2)   | 0.63435(7) | 0.19655(7) | 0.0223(3)*                                                    |
| C6   | 0.4038(2)   | 0.19305(7) | 0.15972(7) | 0.0213(3)                                                     |
| H6   | 0.3738(2)   | 0.13638(7) | 0.16013(7) | 0.0256(3)*                                                    |
| C7   | 0.6051(2)   | 0.22289(8) | 0.19713(7) | 0.0219(3)                                                     |
| H7   | 0.7110(2)   | 0.18649(8) | 0.22407(7) | 0.0263(3)*                                                    |
| C21  | 0.7325(2)   | 0.49885(8) | 0.38178(7) | 0.0237(3)                                                     |
| H21  | 0.8792(2)   | 0.52274(8) | 0.39411(7) | 0.0285(3)*                                                    |
| C14  | 0.3769(2)   | 0.72677(8) | 0.16187(7) | 0.0230(3)                                                     |
| H14  | 0.3004(2)   | 0.76907(8) | 0.18388(7) | 0.0276(3)*                                                    |
| C16  | 0.6621(2)   | 0.68333(8) | 0.08849(7) | 0.0212(3)                                                     |
| H16  | 0.7809(2)   | 0.69603(8) | 0.06041(7) | 0.0254(3)*                                                    |
| C17  | 0.5969(2)   | 0.60292(7) | 0.09589(7) | 0.0186(2)                                                     |
| H17  | 0.6705(2)   | 0.56083(7) | 0.07257(7) | 0.0223(3)*                                                    |
| C15  | 0.5536(2)   | 0.74529(8) | 0.12218(7) | 0.0217(3)                                                     |
| H15  | 0.6004(2)   | 0.80009(8) | 0.11799(7) | 0.0260(3)*                                                    |
| C23  | 0.4258(3)   | 0.41798(8) | 0.41673(7) | 0.0276(3)                                                     |
| H23  | 0.3649(3)   | 0.38535(8) | 0.45228(7) | 0.0331(4)*                                                    |
| C20  | 0.6072(2)   | 0.51182(7) | 0.31190(7) | 0.0183(2)                                                     |
| H20  | 0.6700(2)   | 0.54336(7) | 0.27601(7) | 0.0220(3)*                                                    |
| C24  | 0.2979(2)   | 0.43273(7) | 0.34806(7) | 0.0205(3)                                                     |
| H24  | 0.1474(2)   | 0.41169(7) | 0.33721(7) | 0.0246(3)*                                                    |
| C11  | 0.35390(19) | 0.49572(7) | 0.14733(6) | 0.0141(2)                                                     |
| C22  | 0.6432(2)   | 0.45080(9) | 0.43381(7) | 0.0282(3)                                                     |
| H22  | 0.7310(2)   | 0.44050(9) | 0.48097(7) | 0.0339(4)*                                                    |
| H2   | 0.678(3)    | 0.4602(10) | 0.1758(9)  | 0.025(4)*                                                     |

planes described by the atoms N1–C11–N2 and C12–C11–C18 are almost orthogonal with an intersection angle of 89.8°. The crystal architecture is best viewed along the *a*-axis, exhibiting a head-to-tail arrangement of naphthyl-phenyl layers with interlocking benzoyl rings from the adjacent layer, maintained by a single hydrogen bond, N–H···O = C, and multiple weak Ar···H, Ar···Ar, C = O···H–C contacts.

**Acknowledgements:** The author acknowledges the support of the University of KwaZulu-Natal (Pietermaritzburg) and would like to thank Professor Q. Q. Munro and Dr M. P. Akerman for general advice on matters of crystallography.

## References

1. Bruker. APEX2 and SAINT. Bruker AXS Inc., Madison, Wisconsin, USA, (2012).
2. Sheldrick, G. M.: A short history of SHELX. *Acta Crystallogr. A* **64** (2008) 112–122.
3. Bourhis, L. J.; Dolomanov, O. V.; Gildea, R. J.; Howard, J. A. K.; Puschmann, H.: The anatomy of a comprehensive constrained, restrained refinement program for the modern computing environment – Olex2 dissected. *Acta Crystallogr. A* **71** (2015) 59–75.
4. Dolomanov, O. V.; Bourhis, L. J.; Gildea, R. J.; Howard, J. A. K.; Puschmann, H.: OLEX2: a complete structure solution, refinement and analysis program. *J. Appl. Cryst.* **42** (2009) 339–341.
5. Westrip, S.: publCIF: software for editing, validating and formatting crystallographic information files. *J. Appl. Cryst.* **43** (2010) 920–925.
6. Jeffreys, R. A.: Perimidine Dyes and Intermediates. *J. Chem. Soc.* (1955) 2394–2397.
7. Minkin, V. I.; Zhdanov, Y. A.; Sadekov, I. D.; Raevskii, O. A.; Garnovskii, A. D.: Structure and properties of perimidine and its derivatives. *Chem. Heterocycl. Compd.* **3** (1967) 855–860.
8. Pozharskii, A. F.; Dal'nikovskaya, V. V.: Perimidines. *Russ. Chem. Rev.* **50** (1981) 1559–1600.
9. Herbert, J. M.; Woodgate, P. D.; Denny, W. A.: Synthesis, DNA binding properties, and biological activity of perimidines designed as “minimal” DNA-intercalating agents. *J. Med. Chem.* **30** (1987) 2081–2086.
10. Farghaly, T. A.; Abdallah, M. A.; Muhammad, Z. A.: New 2-heterocyclic perimidines: synthesis and antimicrobial activity. *Res. Chem. Intermed.* **41** (2015) 3937–3947.
11. Maloney, S.; Slawin, A. M. Z.; Woolins, J. D.: 2,2-Dimethyl-2,3-dihydro-1*H*-perimidine. *Acta Crystallogr. E* **69** (2013) o246.
12. Cucciolito, M. E.; Panunzi, B.; Ruffo, F.; Tuzi, A.: Naphthalene-1,8-diamine-2-(pyrimidin-2-yl)-1*H*-perimidine (2/1). *Acta Crystallogr. E* **69** (2013) o1133–o1134.
13. Bello, K. A.; Corns, S. N.; Griffiths, J.: Near-infrared-absorbing squarine dyes containing 2,3-dihydroperimidine terminal groups. *Chem. Commun.* (1993) 452–454.
14. Shaabani, A.; Maleki, A.: Green and efficient synthesis of quinoxaline derivatives via ceric ammonium nitrate promoted and *in situ* aerobic oxidation of  $\alpha$ -hydroxy ketones and  $\alpha$ -keto oximes in aqueous media. *Chem. Pharm. Bull.* **56** (2008) 79–81.
15. Lagrange, A.; Mignon, M.: Oxidative hair dye composition comprising perimidine and other amine coupler in a medium rich in fatty substances. Patent FR2983714A1, 2013.
16. Lagrange, A.; Mignon, M.: Oxidative hair dye composition comprising perimidine and other amine coupler in a medium rich in fatty substances. Patent WO2013087631A2, 2013.