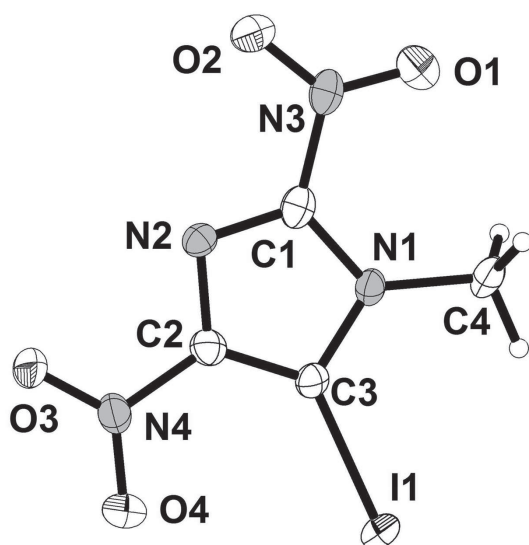


Chen Lizhen, Lian Pengbao and Wang Jianlong*

The crystal structure of 1-methyl-2,4-dinitro-5-iodoimidazole, $C_4H_3IN_4O_4$



DOI 10.1515/ncrs-2016-0209

Received July 7, 2016; accepted December 2, 2016; available online January 11, 2017

Abstract

$C_4H_3IN_4O_4$, orthorhombic, *Pbca* (no. 61), $a = 9.9513(6)$ Å, $b = 11.3365(4)$ Å, $c = 14.3295(12)$ Å, $V = 1616.55(17)$ Å³, $Z = 8$, $R_{gt}(F) = 0.0279$, $wR_{ref}(F^2) = 0.0684$, $T = 105.6$ K.

CCDC no.: 1520552

The asymmetric unit of the title crystal structure is shown in the figure. Tables 1 and 2 contain details on crystal structure and measurement conditions and a list of the atoms including atomic coordinates and displacement parameters.

Source of material

The title compound was prepared by nitrifying 1-menthyl-2,4,5-triiodoimidazole. It was recrystallized from acetone solution at room temperature to give colorless crystals suitable for single-crystal X-ray diffraction.

*Corresponding author: Wang Jianlong, School of Chemical Engineering and Environment, North University of China, Taiyuan 030051, Shanxi Province, P. R. China, e-mail: 619379961@qq.com
Chen Lizhen and Lian Pengbao: School of Chemical Engineering and Environment, North University of China, Taiyuan 030051, Shanxi Province, P. R. China

Table 1: Data collection and handling.

Crystal:	Colourless block
Size:	0.30 × 0.25 × 0.20 mm
Wavelength:	Mo $K\alpha$ radiation (0.71073 Å)
μ :	39.5 cm ⁻¹
Diffractometer, scan mode:	Xcalibur, ω -scans
$2\theta_{max}$, completeness:	52°, >99%
$N(hkl)_{measured}$, $N(hkl)_{unique}$, R_{int} :	4081, 1585, 0.019
Criterion for I_{obs} , $N(hkl)_{gt}$:	$I_{obs} > 2 \sigma(I_{obs})$, 1465
$N(param)_{refined}$:	119
Programs:	CrysAlis ^{PRO} [8], SHELX [9], OLEX2 [10]

Table 2: Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters (Å²).

Atom	x	y	z	U_{iso}^*/U_{eq}
I1	0.07674(3)	0.75983(2)	0.287870(18)	0.01716(12)
N1	0.1556(3)	0.9159(3)	0.1234(2)	0.0126(7)
C4	0.2290(4)	0.9996(3)	0.1853(3)	0.0193(9)
H4A	0.1875	1.0778	0.1811	0.029*
H4B	0.3231	1.0046	0.1656	0.029*
H4C	0.2247	0.9715	0.2499	0.029*
N4	-0.0396(3)	0.6637(3)	0.0563(2)	0.0156(7)
C3	0.0916(4)	0.8142(3)	0.1503(3)	0.0121(8)
C2	0.0405(4)	0.7684(3)	0.0682(3)	0.0140(8)
O4	-0.0608(3)	0.6024(3)	0.1253(2)	0.0261(7)
O3	-0.0825(3)	0.6424(3)	-0.02254(19)	0.0208(7)
O2	0.1504(3)	1.0180(2)	-0.11046(19)	0.0180(6)
O1	0.2522(3)	1.0967(2)	0.0084(2)	0.0214(6)
N2	0.0690(3)	0.8360(3)	-0.0068(2)	0.0120(7)
N3	0.1866(3)	1.0214(3)	-0.0291(2)	0.0152(5)
C1	0.1378(4)	0.9227(3)	0.0290(3)	0.0142(6)

Experimental details

All H atoms were positioned geometrically and treated as riding, with C–H bond lengths constrained to 0.98 Å and $U_{iso}(H) = 1.5U_{eq}(C)$. The diff. electron densities of 1.571 and -0.840 are caused by a partial disorder of I/NO₂.

Discussion

Polynitroimidazole systems have been used as antifungal, antibacterial, antiviral, antitumor drugs and their intermediates [1–3]. Recently, these so called “high energy density materials” have attracted renewed attention in conjunction with their favorable detonation performance [4, 5]. As a promising

candidate, 1-methyl-2,4,5-trinitroimidazole was synthesized by the nitration of 1-methyl-2,4,5-triiodoimidazole [6, 7]. The result showed two iodine atoms were substituted by nitro groups, while the position of another iodine could not be confirmed. In order to determine the molecular structure, the single crystal was cultured and X-ray diffraction experiment is completed.

As shown in the figure, the asymmetric unit of the title compound consists of one molecule. The imidazole ring is almost planar, the two nitro groups are twisted with the imidazole planar, making dihedral angles of 6.0 (N3/O1,O2) and 3.6 (N4/O3,O4)°, respectively. All bond lengths and angles are in the expected ranges.

Acknowledgements: We thank the Center of Testing and Analysis, Beijing University of Chemical Technology, for support.

References

1. Larina, L.; Lopyrev, V.: Nitroazoles: synthesis, structure and application, Springer Verlag, Heidelberg, New York, 2009.
2. Hofmann, K.: Imidazole and Its Derivatives. Part I, John Wiley & Sons, New York 1953.
3. Breccia, A.; Cavalleri, B.; Adams, G. E.: Nitroimidazoles. Chemistry, Pharmacology and Clinical Applications, Vol. 42, Plenum, New York 1982.
4. Cho, R. J.; Kim, J. K.; Cho, G. S.; Kim, J. K.: Synthesis and characterization of 1-methyl-2,4,5-trinitroimidazole (MTNI). *J. Heterocyclic Chem.* **142** (2001) 192–197.
5. Ravi, P.; Tewari, S. P.: Microwave-assisted synthesis of 1-methyl-2,4,5-trinitroimidazole. *Propellants Explos. Pyrotech.* **37** (2012) 544–548.
6. Jadhav, H. S.; Talawar, M. B.; Sivabalab, R.; Dhavalea, D. D.; Asthanab, S. N.; Krishnamurthya, V. N.: Synthesis, characterization and thermolysis studies on new derivatives of 2,4,5-trinitroimidazoles: Potential insensitive high energy materials. *J. Hazard. Mater.* **143** (2007) 611–619.
7. Novikov, S. S.; Khmel'nitskii, L. I.; Lebedev, O. V.; Epishina, L. V.; Sevost'yanova, V. V.: The nitration of iodoimidazoles. *Chem. Heterocycl. Compd.* **6** (1970) 614–618.
8. Agilent Technologies: CrysAlis^{PRO} Software system, version 1.171.35.15, Agilent Technologies UK Ltd, Oxford, UK, 2011.
9. Sheldrick, G. M.: A short history of SHELX. *Acta Crystallogr.* **A64** (2008) 112–122.
10. Dolomanov, O. V.; Bourhis, L. J.; Gildea, R. J.; Howard, J. A. K.; Puschmann, H.: OLEX2: a complete structure solution, refinement and analysis program. *J. Appl. Cryst.* **42** (2009) 339–341.